

ACCELERATED CHEMISTRY MOLECULAR SHAPE

Name: _____
Period: _____

Part 1: For each of the molecules listed complete the table.

	Lewis Structure	Molecular Image (Drawing)
SO_3^{2-}		
BCl_3		
CS_2		
PF_3		
CBr_4		
NO_3^-		

Part II: Fill in the table for each molecule (Use table 1 to help)

	Substance Name	Name of Shape	Angle	Bond Polarity	Molecular Polarity	Hybrid
SO_3^{2-}	Sulfite	Trigonal Pyramidal	107.5°	polar	polar	sp^3
BCl_3	Boron Trichloride	Trigonal planar	120°	polar	Non polar	sp^2
CS_2	Carbon Disulfide	Linear	180°	polar	Nonpolar	sp
PF_3	Phosphorus Trifluoride	Trigonal Pyramidal	107.5°	polar	polar	sp^3
CBr_4	Carbon Tetra bromide	Tetrahedral	109.5°	polar	Non polar	sp^3
NO_3^-	Nitrate	Trigonal Planar	120°	polar	Non polar	sp^2